**Solutions-Acid/Base Review from past units WS**

Graded for correctness

Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Block:\_\_\_\_\_\_\_\_

1. Organize the following words into a flowchart and describe each, homogeneous, heterogeneous, solution, solute, & solvent.

Mixtures

1. Which of the following will conduct an electric current? Why?
   1. K2SO4 (s)
   2. PCl3 (l)
   3. (NH4)2S (aq)
2. Describe how to make a mixture. (include the words solute and solvent)
3. How could one speed up the dissolving (solvation) process?
4. Is the dissolving process (solvation) a physical change or a chemical change? Explain.
5. Explain what it is meant by “likes dissolve likes”.
6. Which of the following will dissolve in water? Why? (may need to do the Lewis structure)
   1. Carbon tetrafluoride
   2. Ammonia (NH3)
   3. Magnesium chloride
7. List and describe at least 4 different physical separation techniques.
8. Draw a pH scale include numeric values and label the acid section, base section, & neutral section.
9. Name or determine the formula for the following:
   1. Hydrochloric acid
   2. N2O
   3. Calcium fluoride
   4. H2SO4
   5. Lead(II) hydroxide