

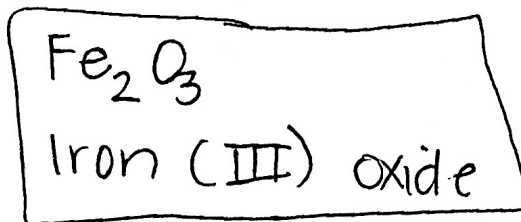
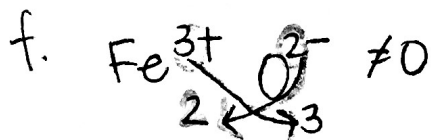
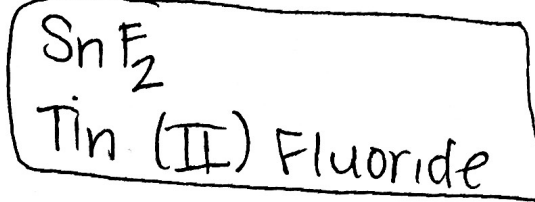
Pg 223 practice #1

- a. potassium & iodine:  $\overset{+1}{\text{K}} \overset{-1}{\text{I}} \rightarrow \boxed{\text{KI}}$
- b. magnesium & chlorine:  $\overset{+2}{\text{Mg}} \overset{-1}{\text{Cl}}_2 \rightarrow \boxed{\text{MgCl}_2}$
- c. sodium & sulfur:  $\overset{+1}{\text{Na}}_2 \overset{-2}{\text{S}} \neq 0 \rightarrow \boxed{\text{Na}_2\text{S}}$
- d. aluminum & sulfur:  $\overset{+3}{\text{Al}}_2 \overset{-2}{\text{S}}_3 \neq 0 \rightarrow \boxed{\text{Al}_2\text{S}_3}$
- e. aluminum & nitrogen:  $\overset{+3}{\text{Al}} \overset{-3}{\text{N}} = 0 \rightarrow \boxed{\text{AlN}}$

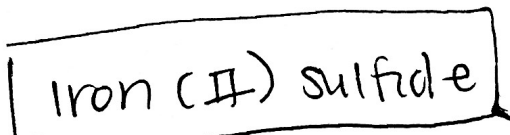
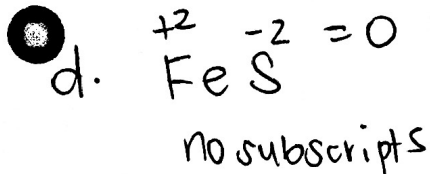
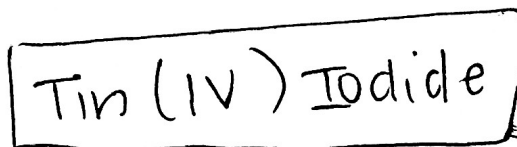
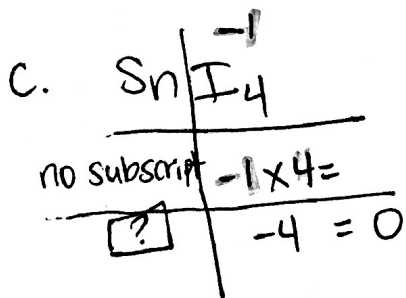
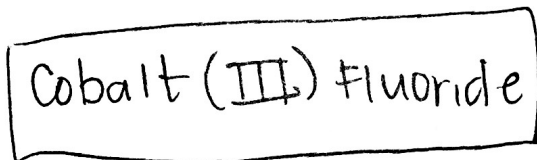
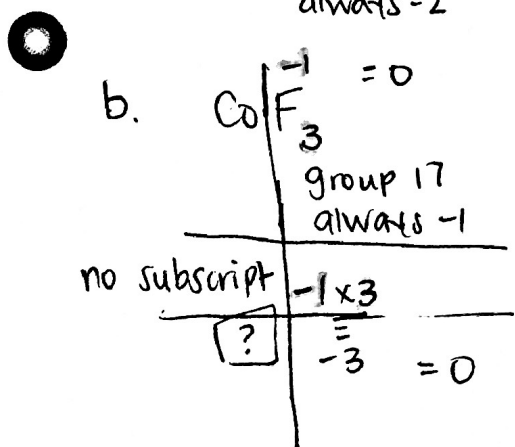
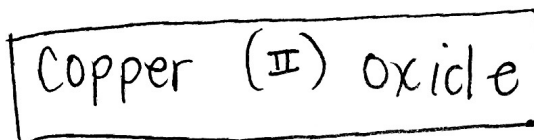
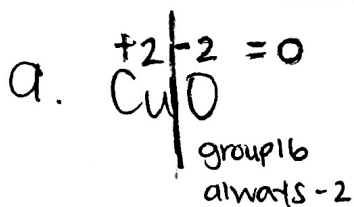
225 practice #1

- a.  $\overset{+2}{\text{Cu}} \overset{-1}{\text{Br}}_2 \neq 0$  Copper (II) bromide  
 $\text{CuBr}_2$
- b.  $\overset{+2}{\text{Fe}} \overset{-2}{\text{O}} = 0$   $\text{FeO}$   
Iron (II) oxide
- c.  $\overset{+2}{\text{Pb}} \overset{-1}{\text{Cl}}_2 \neq 0$   $\text{PbCl}_2$   
Lead (II) chloride
- d.  $\overset{+2}{\text{Hg}} \overset{-2}{\text{S}} = 0$   $\text{HgS}$   
Mercury (II) sulfide

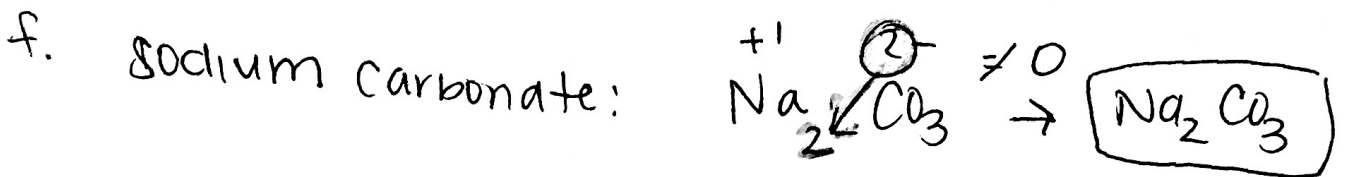
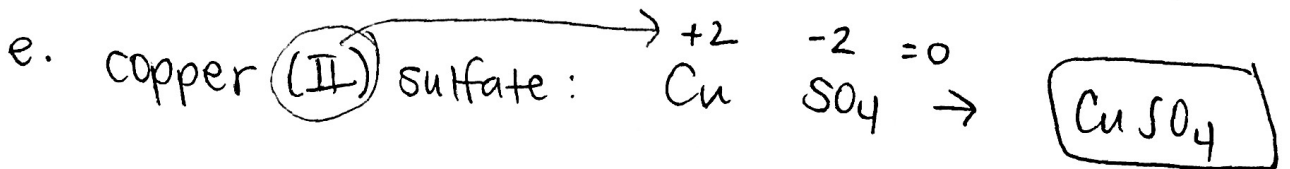
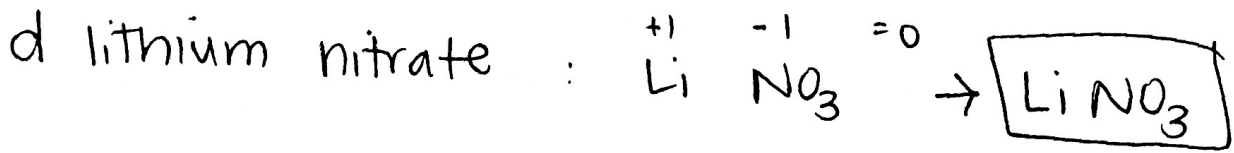
Practice #1 pg 225 cont.



pg 225 practice #2



pg 227 practice #1 D-H



pg 227 practice #2

