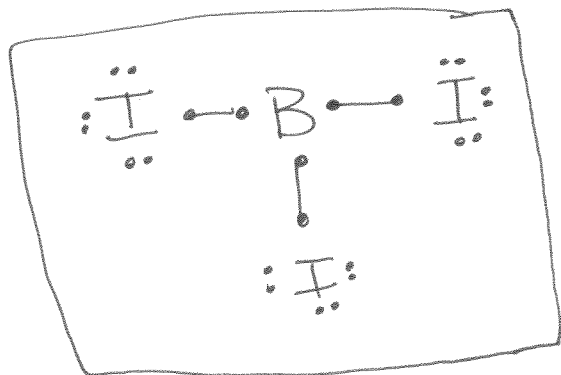


Lewis Structure Homework Answers

1) BI_3 (remember Boron only needs 6 ^{valence} e^-)

B = 3 valence e^- (needs 3 more valence e^-)

I = 7 valence e^- (needs 1 more valence e^-)

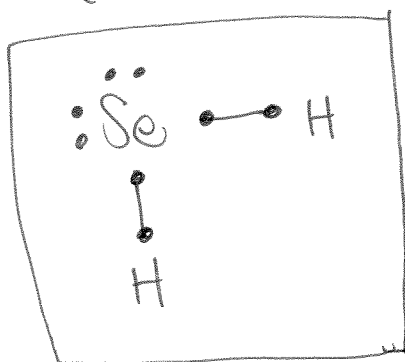


remember,
the element
that needs
the most e^-
goes in the
center

2) SeH_2

Se = 6 valence e^- (needs 2 more e^-)

H = 1 valence e^- (needs 1 more e^-)

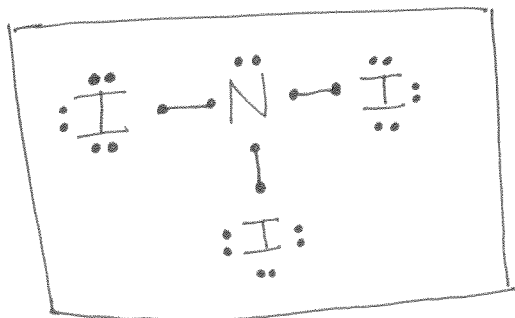


remember
Hydrogen only
needs 2 valence e^-



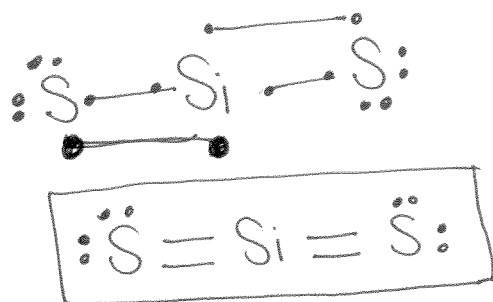
N = 5 valence e^- (needs 3 more e^-)

I = 7 valence e^- (needs 1 more e^-)



Si = 4 valence e^- (needs 4 more e^-)

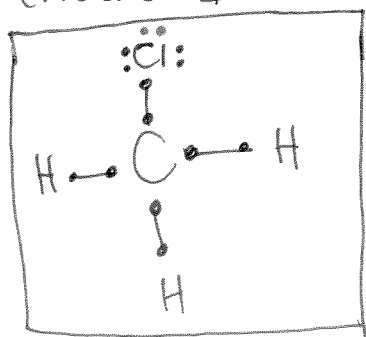
S = 6 valence e^- (needs 2 more e^-)



C = 4 valence e^- (needs 4 more e^-)

H = 1 valence e^- (needs 1 more e^-)

Cl = 7 valence e^- (needs 1 more e^-)



It does not matter where you put the Chlorine