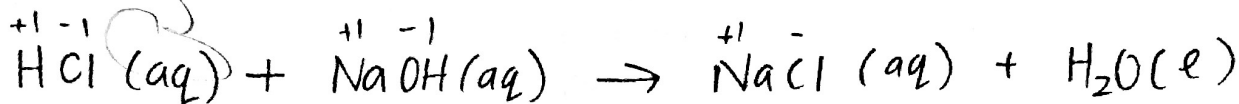
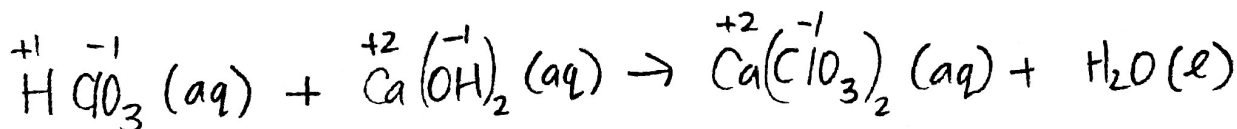


For the following neutralization reactions, write the word equations as chemical formula equations or write the chemical formula equations as word equations. Use table 2 page 266.

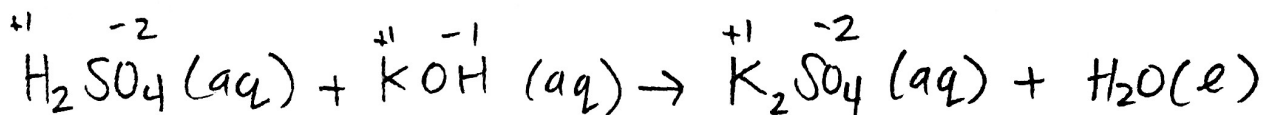
1. Aqueous hydrochloric acid reacts with aqueous sodium hydroxide to produce aqueous sodium chloride and water.



2. Aqueous chloric acid reacts with aqueous calcium hydroxide to produce aqueous calcium chlorate and water.



3. Aqueous sulfuric acid reacts with aqueous potassium hydroxide to produce aqueous potassium sulfate and water.



aqueous nitric acid reacts with aqueous rubidium hydroxide to produce aqueous rubidium nitrate and water



aqueous hydrobromic acid reacts with aqueous strontium hydroxide to produce ~~two~~ aqueous strontium bromide and water

NAMING ACIDS

- **H**Element = hydroelementic acid
ex. HCl = hydrogen + chlorine = hydrochloric acid
- **H**Polyatomic ion ending in **ate** = polyatomic ionic acid
ex. HClO₃ = hydrogen + chlorate = chloric acid
- **H**Polyatomic ion ending in **ite** = polyatomic ionous acid
ex. HClO₂ = hydrogen + chlorite = chlorous acid

Determine the name for each acid below:

- | | |
|---|---|
| 1. H ₂ CO ₃ <u>carbonic</u> acid | 7. H ₃ PO ₄ <u>phosphoric</u> acid |
| 2. H ₂ SO ₄ <u>sulfuric</u> acid | 8. H ₃ PO ₃ <u>phosphorous</u> acid |
| 3. H ₂ SO ₃ <u>sulfurous</u> acid | 9. H ₃ P <u>hydro phosphoric</u> acid |
| 4. H ₂ S <u>hydrosulfuric</u> acid | 10. HNO ₃ <u>nitric</u> acid |
| 5. HC ₂ H ₃ O ₂ <u>acetic</u> acid | 11. HNO ₂ <u>nitrous</u> acid |
| 6. H ₂ O <u>water</u> | 12. H ₃ N <u>hydronitric</u> acid |

Determine the chemical formula for the following.

- | | |
|--|--|
| 1. Hydriodic acid $\begin{matrix} + & - \\ \text{H} & \text{I} \end{matrix}$ | 6. Silver sulfite $\begin{matrix} +1 & -2 \\ \text{Ag}_2 & \text{SO}_3 \end{matrix}$ |
| 2. Sodium iodide $\begin{matrix} +1 & -1 \\ \text{Na} & \text{I} \end{matrix}$ | 7. Hypochlorous acid $\begin{matrix} +1 & -1 \\ \text{H} & \text{ClO} \end{matrix}$ |
| 3. Iron (II) acetate $\begin{matrix} +2 & -1 \\ \text{Fe} & (\text{C}_2\text{H}_3\text{O}_2)_2 \end{matrix}$ | 8. Aluminum chloride $\begin{matrix} +3 & -1 \\ \text{Al} & \text{Cl}_3 \end{matrix}$ |
| 4. Perchloric acid $\begin{matrix} +1 & -1 \\ \text{H} & \text{ClO}_4 \end{matrix}$ | 9. Hydrobromic acid $\begin{matrix} +1 & -1 \\ \text{H} & \text{Br} \end{matrix}$ |
| 5. Calcium chlorite $\begin{matrix} +2 & -1 \\ \text{Ca} & (\text{ClO}_2)_2 \end{matrix}$ | 10. Lead (IV) bromide $\begin{matrix} +4 & -1 \\ \text{Pb} & \text{Br}_4 \end{matrix}$ |