**Acid & Bases: Chap 14 & 15 (pg 467-476)**

1. **The Properties of Acids are:**

|  |  |  |
| --- | --- | --- |
| **a. Sulfuric Acid** | **c.HNO3** | **e. Acetic Acid** |
| **b. H2CO3** | **d. HCl** | **f. hydrobromic acid** |

1. **Name or Determine the formula for the following Acids:**
2. **The Properties of Bases are:**
3. **An Arrhenius Acid is a compound that produces \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions, \_\_\_\_\_\_\_, in aqueous solution. Aqueous means that a compound is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in water.**
4. **Strong Acids are \_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrolytes.**
5. **An Arrhenius Base is a compound that produced \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions, in aqueous solution.**
6. **Strong Bases are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ electrolytes.**
7. **Bronsted-Lowry acid is a molecule or ion that is a \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_. Because \_\_\_\_\_\_ is a \_\_\_\_\_\_\_\_\_, all acids defined by Arrhenius donate \_\_\_\_\_\_\_\_\_.**
8. **Bronsted-Lowry base is a molecule or ion that is a \_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_.**
9. **Copy the reaction between hydrochloric acid (HCl) & ammonia (NH3):**
   1. **In this reaction who is the Bronsted-Lowry acid and why?**

* 1. **In this reaction who is the Bronsted-Lowry base and why?**

1. **Give an example of a strong acid-strong base neutralization reaction.**
2. **Copy figure 3 on page 503.**